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this set (21) Stoichiometry . The calculation of quantities in chemical reactions is a subject of chemistry. Mole ratio. A conversion factor derived from coefficients of a balanced chemical equation ...

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Everyday Stoichiometry. You have learned about chemical equations and the techniques used in order to balance them. Chemists use balanced equations to allow them to manipulate chemical reactions in a quantitative manner. Before we look at a chemical reaction, let's consider the equation for the ideal ham sandwich. Figure 12.1.1: The ideal ham sandwich. Our ham sandwich is composed of 2 slices ...

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Name CHAPTER STUDY GUIDE Date Class Stoichiometry Section 11.1 What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation.

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Answers 1. Because there is more than enough of that material and its amount will not affect the mole ratios of other components of the reaction. 2. 3 moles H/1 mole N. 3. There are 6 H atoms and 2 C atoms in the molecule, so there are 6H/2C or a mole ratio of H:C of 3:1. 12.3 Mass-Mole and Mole-Mass Stoichiometry Practice

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Science / Chapter 12 - stoichiometry (handouts)

1 Chapter 12 Multiple Choice Identify the letter of the choice that best completes the statement or answers the question. ____ 1. The first step in most stoichiometry problems is to ____.

- a. add the coefficients of the reagents
- c. convert given quantities to volumes
- b. convert given quantities to moles
- d. convert given quantities to masses

____ 2.

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